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 $\text{H}_2\text{O (g)} + \text{H}_2 \text{ (g)} + \frac{1}{2} \text{ O}_2 \text{ (g)}$   $\Delta H = 71.2 \text{ kJ}$   
 $\text{C}_2\text{H}_6 \text{ (g)} + \frac{7}{2} \text{ O}_2 \text{ (g)} \rightarrow 2 \text{ CO}_2 \text{ (g)} + 3 \text{ H}_2\text{O (g)}$   $\Delta H = 283 \text{ kJ}$   
g 1  $\text{C}_2\text{H}_6 \text{ (g)}$

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 $\text{C}_2\text{H}_4 \text{ (g)} + 3 \text{ O}_2 \text{ (g)} \rightarrow 2 \text{ CO}_2 \text{ (g)} + 2 \text{ H}_2\text{O (l)}$   $\Delta H = 1411 \text{ kJ}$   
 $\text{C}_2\text{H}_6 \text{ (g)} + \frac{7}{2} \text{ O}_2 \text{ (g)} \rightarrow 2 \text{ CO}_2 \text{ (g)} + 3 \text{ H}_2\text{O (l)}$   $\Delta H = 1560 \text{ kJ}$   
mole  $\text{H}_2 \text{ (g)} + \frac{1}{2} \text{ O}_2 \text{ (g)} \rightarrow \text{H}_2\text{O (l)}$   $\Delta H = 285 \text{ kJ}$  mole

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web hess s laws abstract the following experiment was designed using hess s law and simple calorimetry to determine the enthalpy for naoh s hess s law says that the overall change in enthalpy for the reaction is the sum of all changes regardless of the different stages or phases of a reaction

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web hess s law states that if a reaction is the sum of two or more other reactions the  $\Delta h$  for the overall process must be the sum of the  $\Delta h$  values of the constituent reactions enthalpy change  $\Delta h$  is independent of the path that a reaction follows to move from reactants

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web this activity provides a demonstration of hess law using three reactions the solubility naoh in water the solubility naoh in hcl and the reaction of a solution of hcl and a solution of naoh

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web hess s law worksheet answers 1 calculate h for the reaction c2h4 g h2 g c2h6 g from the following data c2h4 g 3 o2 g 2 co2 g 2 h2o l h 1411 kj c2h6 g 3½ o2 g 2 co2 g 3 h2o l h 1560 kj

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web mar 13 2023 be able to use hess s law to calculate reaction enthalpies first law of thermodynamics energy is the ability to do work or transfer heat work is the transfer of energy from one body to another in a sense work is energy in the process of transfer

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web may 8 2015 hess law lab by maya parks partners ben seufert kelsea floyd convert joules to kj in your final answer multiply the mass by the change in temperature and the c given trial 1 use hess s law and the following equations and  $\Delta h$  values to determine the heat of reaction for the reaction c 2 h 4

**experiment 9 enthalpy of reaction hess s law** - Jul 14 2023

web thus for reaction 9 2 in which 1 mole of h3o and 1 mole of oh combine to form 2 moles of h2o h 55 8 kj as shown note that the enthalpy change h hproducts hreactants is positive if heat is absorbed that is if hproducts

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**reaction studocu** - Nov 06 2022

web hess s law enthalpy of reaction and hess s law abstract  
the goal of this lab is to perform experiments and practice  
hess s law by determining enthalpy change in chemical  
reactions in this lab three experiments were conducted in  
order to test hess s law the experiments required observing  
and collecting data on three different reactions

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07 2022

web objective the purpose of the lab is to combine two  
reactants in the calorimeter and measure the heats of

reaction in solution eventually allowing an individual to  
prove or disprove hess s law introduction thermodynamics is  
the study of energy exchanged between the system and the  
surroundings

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Mar 10 2023

web 3 12 hess s law in this experiment you will measure the  
amount of heat released in these three related exothermic  
reactions 1  $\text{NaOH (s)} + \text{NaOH (aq)} \rightarrow \text{H}_2\text{O (l)}$  2  $\text{NaOH (s)} + \text{HCl (aq)} \rightarrow \text{H}_2\text{O (l)}$   
3  $\text{NaOH (aq)} + \text{HCl (aq)} \rightarrow \text{H}_2\text{O (l)}$